

**AD-A202 078**

**DTIC FILE COPY**

(4)

OFFICE OF NAVAL RESEARCH

Contract NOOO14-80-K-0852

R&T Code \_\_\_\_\_

Technical Report No. 45

Understanding Core Level Decay Processes  
In the High-Temperature Superconductors

By

D. E. Ramaker, N. H. Turner and F. L. Hutson

Prepared for Publication

in the

Physical Review B

George Washington University  
Department of Chemistry  
Washington, D.C. 20052

December, 1988

Reproduction in whole or in part is permitted for  
any purpose of the United States Government

This document has been approved for public release  
and sale; its distribution is unlimited.

**DTIC  
ELECTED  
S DEC 19 1988  
or  
H**

ADA202078

SECURITY CLASSIFICATION OF THIS PAGE

REPORT DOCUMENTATION PAGE

1a. REPORT SECURITY CLASSIFICATION Unclassified		1b. RESTRICTIVE MARKINGS										
2a. SECURITY CLASSIFICATION AUTHORITY		3. DISTRIBUTION/AVAILABILITY OF REPORT Approved for Public Release, distribution Unlimited.										
2b. DECLASSIFICATION/DOWNGRADING SCHEDULE												
4. PERFORMING ORGANIZATION REPORT NUMBER(S) Technical Report # 45		5. MONITORING ORGANIZATION REPORT NUMBER(S)										
6a. NAME OF PERFORMING ORGANIZATION Dept. of Chemistry George Washington Univ.	6b. OFFICE SYMBOL (If applicable)	7a. NAME OF MONITORING ORGANIZATION Office of Naval Research (Code 413)										
6c. ADDRESS (City, State, and ZIP Code) Washington, D.C. 20052		7b. ADDRESS (City, State, and ZIP Code) Chemistry Program 800 N. Quincy Street Arlington, VA 22217										
8a. NAME OF FUNDING/SPONSORING ORGANIZATION Office of Naval Research	8b. OFFICE SYMBOL (If applicable)	9. PROCUREMENT INSTRUMENT IDENTIFICATION NUMBER Contract N00014-80-K-0852										
8c. ADDRESS (City, State, and ZIP Code) Chemistry Program 800 North QUINCY, Arlington, VA 22217		10. SOURCE OF FUNDING NUMBERS <table border="1"> <tr> <td>PROGRAM ELEMENT NO. 61153 N</td> <td>PROJECT NO. 013-08-01</td> <td>TASK NO. PP</td> <td>WORK UNIT ACCESSION NO. NR 056-681</td> </tr> </table>		PROGRAM ELEMENT NO. 61153 N	PROJECT NO. 013-08-01	TASK NO. PP	WORK UNIT ACCESSION NO. NR 056-681					
PROGRAM ELEMENT NO. 61153 N	PROJECT NO. 013-08-01	TASK NO. PP	WORK UNIT ACCESSION NO. NR 056-681									
11. TITLE (Include Security Classification) Understanding Core Level Decay Processes in the High-Temperature Superconductors (Uncl.)												
12. PERSONAL AUTHOR(S) D. E. Ramaker, N. H. Turner, and F. L. Hutson												
13a. TYPE OF REPORT Interim Technical	13b. TIME COVERED FROM _____ TO _____	14. DATE OF REPORT (Year, Month, Day) December 1988	15. PAGE COUNT 10									
16. SUPPLEMENTARY NOTATION Prepared for publication in Physical Review B												
17. COSATI CODES <table border="1"> <tr> <th>FIELD</th> <th>GROUP</th> <th>SUB-GROUP</th> </tr> <tr> <td> </td> <td> </td> <td> </td> </tr> <tr> <td> </td> <td> </td> <td> </td> </tr> </table>		FIELD	GROUP	SUB-GROUP							18. SUBJECT TERMS (Continue on reverse if necessary and identify by block number) Superconductivity, Auger Spectroscopy, X-ray Emission Spectroscopy, Hubbard Model, <i>Copy 2</i> , <i>copy 3</i> , <i>copy 4</i> , <i>copy 5</i>	
FIELD	GROUP	SUB-GROUP										
19. ABSTRACT (Continue on reverse if necessary and identify by block number) A highly correlated CuO <sub>7</sub> <sup>7</sup> cluster model is utilized to interpret the core level Auger and X-ray emission decay spectra for YBa <sub>2</sub> Cu <sub>3</sub> O <sub>7-x</sub> and CuO. The evidence indicates that the initial-state shakeup states relax to states of the same symmetry before the core level decay, provided they have a shakeup excitation energy much greater than the core level width.												
20. DISTRIBUTION/AVAILABILITY OF ABSTRACT <input checked="" type="checkbox"/> UNCLASSIFIED/UNLIMITED <input checked="" type="checkbox"/> SAME AS RPT. <input type="checkbox"/> DTIC USERS		21. ABSTRACT SECURITY CLASSIFICATION Unclassified										
22a. NAME OF RESPONSIBLE INDIVIDUAL Dr. David L. Nelson		22b. TELEPHONE (Include Area Code) (202) 696-4410	22c. OFFICE SYMBOL									

DD FORM 1473, 84 MAR

83 APP edition may be used until exhausted.  
All other editions are obsolete.

SECURITY CLASSIFICATION OF THIS PAGE  
Unclassified

Previously reported [1,2] core level Auger (AES) and x-ray emission (XES) data are interpreted within a highly correlated CuO<sub>n</sub> cluster model for the high-temperature superconductors (HTSC's), YBa<sub>2</sub>Cu<sub>3</sub>O<sub>7-x</sub> and La<sub>2</sub>Ba<sub>2</sub>CuO<sub>6</sub> (herein referred as 123 and La). The La<sub>2</sub>M<sub>3</sub>V Auger lineshape is reported here for the first time and is interpreted consistently with the La<sub>2</sub>VV lineshape and XES data. This work clearly indicates, contrary to previous reports [1,3], that the initial-core shake-up (ICSU) states do not directly decay, but rather relax to the primary core-state before decay. The XES data dramatically reveal the change of character of the valence band (VB) states between CuO and 123.

The basic electronic structure of the HTSC's can be described by an extended Hubbard model, characterized by the transfer or hopping integral  $t$ , the Cu and O orbital energies  $\epsilon_{\text{Cu}}$  and  $\epsilon_{\text{O}}$ , the core polarization energy  $Q_d$ , the intra-site Coulomb repulsion energies  $U_{dd}$  and  $U_{pp}$ , and the inter-site repulsion energies  $U_{dp}$  and  $U_{pp'}$  (i.e. between neighboring Cu-O and O-O atoms). The extended Hubbard model is most appropriate when the U's are large relative to the band widths [3], i.e. when correlation effects dominate covalent or hybridization effects. A CuO<sub>n</sub><sup>(2n-2)-</sup> cluster model, which is also reasonably valid when  $U \gg t$ , simplifies the model [3]. Both La and CuO contain CuO<sub>6</sub> groups [4], having 4 short and 2 long Cu-O bonds. The 123 HTSC contains CuO<sub>6</sub> and planar CuO<sub>4</sub> groups [4]. The different n may alter the relative intensities of various features, but similar features are present in each case. The different bond lengths may increase the widths of the spectral features, but little else since correlation dominates.

The CuO<sub>n</sub><sup>(2n-2)-</sup> cluster has one hole shared between the Cu 3d and O 2p shells in the ground state, which we term the v (valence) states. The spectroscopic final states reflect multi-hole states, e.g. v<sup>2</sup>, cv (c = core) etc.

We indicate the location of the  $\nu$  holes by d (Cu 3d) or p (O 2p). In the case of two holes on the oxygens, we distinguish two holes on the same O (p $\downarrow$ ), on ortho neighboring O atoms (pp $\downarrow$ ), or on para O atoms (pp $\pm$ ) of the cluster. Furthermore, neighboring pp $\pm$  holes can dimerize [5], so we distinguish between two holes in bonded (pp $\pm$ ) and antibonded (pp $\mp$ ) O pairs. Most of the O atoms actually participate in two CuO<sub>6</sub> clusters. Consistent with previous work [6], we account for this by defining the effective parameter,  $\epsilon_p = \epsilon_{p\downarrow} + U_{p\downarrow}$ , where  $U_{p\downarrow}$  includes the interaction of a hole in an O p orbital with its environment. In general  $U_{p\downarrow}$  will be less than  $U_{d\downarrow}$  due to polarization.

The  $\nu$  states, as reflected by the theoretical DOS [7], can be described as having the Cu-O bonding ( $t_b$ ) and antibonding ( $t_a$ ) orbitals centered at 4 and 0 eV and the nonbonding Cu and O orbitals at 2 eV. The O features each have a width  $2\Gamma = 4$  eV due to the O-O bonding and antibonding character and the Cu-O dispersion. The  $t_b$  and  $t_a$  wavefunctions can be expressed as [3],

$$t_a = d \cos\theta_1 - p \sin\theta_1 \quad (1a)$$

$$t_b = d \sin\theta_1 + p \cos\theta_1, \quad (1b)$$

where  $\theta_1 = 0.5 \tan^{-1}(2t/\Delta)$ . We also define the Cu-O hybridization shift  $\delta_1 = 0.5 \sqrt{(\Delta^2+4t^2)} - \Delta/2$ , which is utilized in Table 1 to give the energies. In this picture, the ground state of an average CuO<sub>6</sub> cluster is located at 1 eV having the energy  $\epsilon_d - \delta_1 + \Gamma/2 = \epsilon_d - \alpha$ , which we use as a reference energy for the excited  $\nu^2$ ,  $\nu^3$ , and  $c\nu$  states reflected in the core level spectra. In CuO, the hybridization shift  $\Gamma$  is smaller, and we shall see below that  $\Delta \approx \epsilon_d - \epsilon_a$  has increased to 1 eV.

Recently [8] we consistently interpreted the VB photoelectron spectra (UPS and XPS). Most of the features in the UPS and XPS are also reflected

Quality Codes	
INSPECTED	Avail and/or Special
2	
A-1	

in the AES and XES, so that we review these assignments here. In 123, the states were assigned as indicated in Table 1 [8,10]. Calculated photoemission intensities, their variation with  $\Delta$ , and photon energy dependencies confirm these assignments [8]. In CuO [9], we have previously assigned a feature at 5.5 eV to pp<sup>z</sup> and pp<sup>y</sup> and at 3 eV to dp. The character switch of state 1 from mostly dp to pp<sup>y</sup> and vice versa for state 2 between CuO and 123 arises because  $\Delta$  decreases from 1 eV in CuO to 0 eV in 123. The reduction in  $\Delta$  as indicated by the UPS data is consistent with the Cu 2p XPS data and with the XES data to be discussed below.

The "shakeup" or "many-particle" features at 9.5, 12.5, and 16 eV have also been assigned as indicated in Table 1 [8]. The pp<sup>z</sup> state has been assigned to the "mystery" feature at 9.5 eV. Such a feature also appears for CuO [9,11] so that this feature is not unique to the HTSC's. This feature cannot arise from the p<sup>z</sup> final state because U<sub>f</sub> is around 12-13 eV, much too large to cause a feature at 9.5 eV. In fact we have found evidence [8] for the existence of the p<sup>z</sup> feature around 16 eV in 123. Finally the d<sup>z</sup> state is known to cause the 12.5 eV feature [12].

Cu 2p and O 1s core level XPS. In order to understand the XES and AES data, we first characterize the initial state, which is reflected directly in the Cu 2p and O 1s XPS data. The primary and satellite features seen in the Cu 2p XPS spectrum for CuO [13] and 123 or La [1,14] are known to arise from the cp and cd states, respectively [3], having the energies given in Table 1. The relative satellite intensity, I(cp)/I(cd) decreases from 0.55 in CuO to 0.37 in 123 as determined from the experimental data [1]. The energy separation, E(cd) - E(cp) increases from 8.7 eV in CuO to 9.2 eV in 123 [1].

The primary (cp) and satellite (cd) wavefunctions can be written similar to eq. (1), with hybridization angle  $\theta_c = 0.5 \tan^{-1}(2t/(\Delta-Q_d))$  [3]. In the

sudden approximation, the intensities are proportional to the overlap between the ground state wavefunction,  $\psi_0$ , and the final states, so that  $I(dp) = \cos^2(\theta_c - \theta_1)$  and  $I(cd) = \sin^2(\theta_c - \theta_1)$  [3]. Thus the satellite intensity increases with change in the hybridization angles between the v and cv states. In the ground v state, the hole is shared equally in the p and d orbitals since  $j_1 = 45^\circ$ , in the primary cv state it is mostly in the p orbital since  $j_C = 78^\circ$ . The changes between CuO and 123 noted above are just that expected for a decrease in  $\Delta$  and reflect an increased covalency in 123 [15].

The large width of the primary cp peak is believed to arise from the mixing with the cd state [3,15]. The cd state has a large width due to the large core-hole, valence-hole interaction, indeed, the satellite actually reveals the cd multiplet structure. Evidence that the primary cp peak width arises from the cd interaction comes from the Cu halide data [3], which show a direct correlation of the primary cp peak width with the satellite cd peak intensity. We do not believe that the primary peak width arises from the O p band width as proposed by others [16].

The O 1s spectra have been reported by many authors; however, it is seriously altered by impurities such as OH<sup>-</sup> and CO<sub>3</sub><sup>2-</sup> on the sample surface [17]. Recent data [18] from single crystal samples of the La material cleaved in-situ are expected to be reasonably free of impurity effects. The cp<sup>\*</sup> and cp<sup>+</sup> states listed in Table 1 are believed to account for the tailing off of the spectra seen in these spectra (this will be positively identified upon examination of the XES data). Consistent with the sudden approximation, the cp state is not seen in the O 1s XPS because now both the v and cv states have similar hybridization angles, i.e. the valence hole is mostly in the d orbital in both cases.

We will find below that the ISSU process, which is responsible for the satellites in the XPS noted above, does not produce satellites in the AES or XES data, because the ISSU states generally "relax" to the primary states of the same symmetry before the core level decay. Such a relaxation is expected when the ISSU excitation energy is larger than the core level width [19].

Previously, vanderLaan et al [3] suggested the intensity of these ISSU states in the XPS should be quantitatively reflected in the intensity of the Auger satellites found in the  $\text{La}_{2}\text{VV}$  lineshapes for the Cu halides. The data do not indicate this however. While  $I(\text{cd})/I(\text{cp})$  increases from 0.45 for  $\text{CuBr}_2$  to 0.8 for  $\text{CuF}_2$ , the Auger satellites do not increase [3]. We previously [1] indicated that a fraction of these ISSU states probably resulted in Auger satellites for the HTSC's, and that this fraction increased with the increasing covalency of the HTSC material. Evidence presented here indicates rather that the ISSU states relax before the core level decay to states of the same symmetry, provided they have a ISSU excitation energy much greater than the core level width. We believe this to be a general result, at least in the  $\text{Cu}^{+2}$  materials.

The  $\text{Cu L}_{2,3}$  and O K XES data. The O K XES data [2] in Figure 1a confirms our assignment of the O XPS, and clearly shows the dependency of the ISSU state relaxation on the excitation energy and symmetry. The principal XPS peak arises from the cd state, and it decays to the dp state since the x-ray emission process is intra-atomic in nature. Therefore the principal O XES peak aligns with the dp feature in the UPS as shown in Fig. 1. The  $\text{cp}^{\circ}$  state does not mix with the primary cd state; therefore, it does not relax before the decay, but decays directly to the  $\text{pp}^{\circ\circ}$  (and perhaps a little also to the  $\text{pp}^{\circ\circ\circ}$ ) state. This accounts for the feature around 6.5 eV in the XES, just 3 eV above the  $\text{pp}^{\circ\circ}$  feature in the UPS. The shift of 3 eV matches the energy

difference between the  $cp^*$  and  $cd$  core hole states. The  $cp^*$  state can mix with the  $cd$  state, therefore it can relax to the  $cd$  state, but it does this slowly because of the small excitation energy of 0.5 eV. Therefore, the  $cp^*$  state decays either directly to the  $pp^*$  state, or relaxes to the  $cd$  state, which then decays to the  $dp$  state. This explains the photon energy dependence seen [2] in the data of Fig. 1a. At high photon energy, the sudden approximation is more valid, creating a larger intensity for the  $cp^*$  state, and consequently a larger  $pp^*$  contribution around 2.5 eV in the XES.

The Cu L<sub>2,3</sub> XES data [2,20] shown in Figure 1b dramatically reveals the switch in character of the 1 and 2  $v^2$  states between CuO and 123. Again, the satellite  $cd$  initial state relaxes to the  $cp$  state before the decay so that the XES reflects primarily the  $dp$  DOS. In CuO the XES spectrum peaks at 3 eV, in 123 it falls around 4.2 eV, very near where we indicated the  $dp$  states fall in the UPS data. The large intensity in the CuO XES extending above the Fermi level is believed to be an experimental artifact [20].

The Cu L<sub>2,3</sub>VV and L<sub>2,3</sub>M<sub>2,3</sub>V Auger data. Comparison of the L<sub>2</sub>VV data for CuO [11] and 123 [1] are shown in Fig. 2. The data reveal features at 7 (the two-center feature), 15, and 19 eV, which we previously [1] attributed to  $dp$ ,  $d^2$ , and  $d^3$  final states, utilizing a  $v^2$  final state model. The  $d^3$  states were attributed to a combination of 3 different processes; 1) initial state shakeoff (ISSO) followed by Auger decay ( $g.s. + h\nu \rightarrow L_2v \rightarrow d^3$ ), 2) Coster-Kronig (CK) decay followed by Auger decay ( $g.s. + h\nu \rightarrow L_2u \rightarrow L_2v \rightarrow d^3$ ), and 3) ISSU followed by Auger decay ( $g.s. + h\nu \rightarrow L_2ve \rightarrow L_2v \rightarrow d^3$ , where  $e$  denotes the excited electron). The ISSO and CK processes accounted for all of the  $d^3$  component in CuO, and the ISSU process was believed, as mentioned above, to account for the increasing  $d^3$  component in La and 123 [1].

We report and interpret here, for the first time, the L<sub>2,3</sub>M<sub>2,3</sub>V Auger

lineshapes for the 123 SC. The sample preparation, treatment, and instrument utilized were described previously [1]. Fig. 2 compares the  $L_{23}M_{23}V$  spectra for CuO ([24] and 123, and identifies the various features. The  $L_{23}M_{23}V$  lineshapes reflect the  $cv^1$  DOS, the main features arising from the cdp final state, and the satellite from the  $d^3p$  state apparently resulting from the similar ISSO, CK, and ISSU processes defined above. However, Fig. 2 reveals a most interesting point; although 123 shows an increased satellite in the  $L_{23}VV$  relative to CuO, it is not increased in the  $L_{23}M_{23}V$ . This indicates strongly that the ISSU process is not responsible for the increased satellite in the  $L_{23}VV$ , because then it should increase the satellite in both 123 lineshapes.

Since only the primary cp core-hole state Auger decays, and this process is also known to be strictly intra-atomic, the  $L_{23}VV$  lineshape in our current  $v^3$  final state model, reflects the  $d^3p$  DOS, as it is distributed among the  $v^3$  states listed in Table 1. Thus the features at 7, 15, and 19 eV arise naturally from the  $dpp^1$ ,  $d^3p$ , and  $dp^2$  final states. The ISSO and CK processes also contribute to the "satellite" contribution at 19 eV just as in CuO. The  $dpp^1$  state does not appear in the  $L_{23}VV$  lineshape because it does not have the same symmetry possessed by all the other  $v^3$  final states and the  $cv$  initial state. The increased "satellite" feature at 19 eV in the HTSC's arises apparently because of increased configuration mixing between the  $d^3p$  and  $dp^2$  states. Its intensity is increased in 123 relative to CuO because the energy separation (before hybridization) between  $d^3p$  and  $dp^2$  has decreased from 3.8 eV in CuO to 2.5 eV in 123. We have indicated this mixing in Table 1 by adding the hybridization shifts  $\delta_z$  to the energy expressions for these two states.

The  $L_{23}M_{23}V$  lineshape reflects the cdp DOS. The mixing of the other states ( $cd^3$ ,  $cpp^1$ ,  $cpp^0$ , and  $cpp^-$ ; the latter three are not listed in Table 1)

with the cdp state is small because of the large energy separations involved. The cp<sup>2</sup> state is close to cdp; however, it falls in between the ^3L and ^1L multiplets of the cdp state. Although it may have some intensity, it surely does not contribute to the CK + SU satellite around 25 eV in either CuO or 123. The exchange splitting (2K) between the 3p and d holes is known to be very large [3], so we include it explicitly in Table 1 to account for the ^1L multiplets.

The O KVV lineshape is severely altered by impurities on the sample surfaces, and no single crystal lineshape data have been reported. The O KVV lineshapes for CuO and Cu<sub>2</sub>O have been reported [11], and they have the primary dp<sup>2</sup> or p<sup>2</sup> features, respectively, around 19 eV. A very small satellite appears around 7 eV in Cu<sub>2</sub>O which we attribute to the pp<sup>2</sup> state. A much larger and broader satellite around 7 to 14 eV in CuO appears, which we attribute to the d<sup>2</sup>p state around 14 eV as well as a smaller amount to the dpp<sup>2</sup> state around 7 eV. Thus the d<sup>2</sup>p and dp<sup>2</sup> states appear in both the Cu L<sub>2,3</sub>VV and O Auger lineshapes for Cu<sup>2+</sup> oxides, except their primary and satellite roles are reversed.

In summary, we have interpreted XES and AES data utilizing a highly correlated CuO<sub>n</sub> cluster model. Both the XES data and the previously interpreted UPS data reveal the reversal in character of the VB states between CuO and the HTSC's. We have also shown that the initial-state shake-up states evident in core level XPS, do not generally produce satellites in the core emission spectra, because they relax to the primary core states of the same symmetry, provided the ISSU excitation energy is greater than the core level width.

TABLE 1 Summary of hole states revealed in the spectroscopic data, and estimated energies using the following optimal values for the Hubbard parameters in eV<sup>a</sup>:

$\delta_1 = 2$        $t_c = 2$        $U_p = 12, 13$        $U_d = 9.5, 10.2$   
 $\delta_2 = 0.5, 0.8$        $c_p = 2, 3$        $U_{pp^*} = 4.5, 4$        $U_{dd} = 1$   
 $\Gamma = 2$        $U_{pd} = 0$        $U_{cp^*} = 2$        $Q_d = 9$   
 $\alpha = 1, 0.5$        $\Delta = 0, 1.$        $K = 4$

State <sup>b</sup>	Energy expression	Calc. E. eV <sup>c</sup>	Exp. E. eV <sup>c</sup>	Remark
<u>G.S. and IPES, v<sup>d</sup></u>				
+ <sub>d</sub> ) d	$c_p - \delta_1 + \Gamma$	0 ± 2	-	heavily mixed
+ <sub>p</sub> ) p	$c_p + \delta_1 + \Gamma$	4 ± 2	-	
<u>UPS and XES, v<sup>e</sup></u>				
1) pp <sup>*</sup>	$c_p + \Delta - \delta_2 + \alpha$	2.5	2.5	heavily mixed
2) dp	$c_p + U_{dp} + \delta_2 + \alpha$	4.5	4.2	
3) pp <sup>*</sup>	$c_p + \Delta + U_{pp^*} - \Gamma + \alpha$	5.5	5.	
4) pp <sup>*</sup>	$c_p + \Delta + U_{pp^*} + \Gamma + \alpha$	9.5	9.5	mystery peak
5) d <sup>2</sup>	$c_d + U_d + \alpha$	12.5	12.5	Cu sat.
6) p <sup>2</sup>	$c_p + \Delta + U_p + \alpha$	15	16	
<u>Cu 2p XPS, cv</u>				
cp	$c_c + \Delta + \alpha$	$c_c + 1$	$E_{1s}$	main
cd	$c_c + Q_d + \alpha$	$c_c + 10$	$E_{1s} + 9.2$	sat.
<u>O 1s XPS, cv</u>				
cd	$c_c + \alpha$	$c_c + 1$	$E_{1s}$	main
cp <sup>*</sup>	$c_c + \Delta + \alpha$	$c_c + 1$	$E_{1s}$	main
cp <sup>*</sup>	$c_c + \Delta + U_{cp^*} + \alpha$	$c_c + 3$	$E_{1s} + 2$ ?	tail
cp	$c_c + \Delta + Q_p + \alpha$	?	?	not obs
<u>Cu L<sub>2,3</sub>VV AES, v<sup>f</sup></u>				
dpp <sup>*</sup>	$2c_p + 2U_{dp} + \alpha$	7	7	2 cent. feature
dpp <sup>*</sup>	$2c_p + U_{pp^*} + 2U_{dp} + \alpha$	11.5	-	no mixing
d <sup>2</sup> p	$c_d + c_p + U_d + 2U_{dp} - \delta_2 + \alpha$	16	15.5	main feature
dp <sup>2</sup>	$2c_p + U_p + 2U_{dp} + \delta_2 + \alpha$	19.5	18-25	sat. feature
<u>Cu L<sub>2,3</sub>M<sub>1,2</sub>V ABS, cv<sup>f</sup></u>				
cdp	$c_c + c_p + Q_d + U_{dp} + K + \alpha$	$c_c + 9$	$E_{1s} + 10$	main, <sup>1</sup> L
		$c_c + 17$	$E_{1s} + 18$	main, <sup>3</sup> L
cp <sup>*</sup>	$c_c + c_p + \Delta + U_p + \alpha$	$c_c + 15$	-	not observed
cd <sup>2</sup>	$c_c + c_d + U_d + 2Q_d + \alpha$	$c_c + 30.5$	-	not observed

<sup>a</sup>Parameters for 123 indicated first, those for CuO second.

<sup>b</sup>The dominant character in the hybridized states is given.

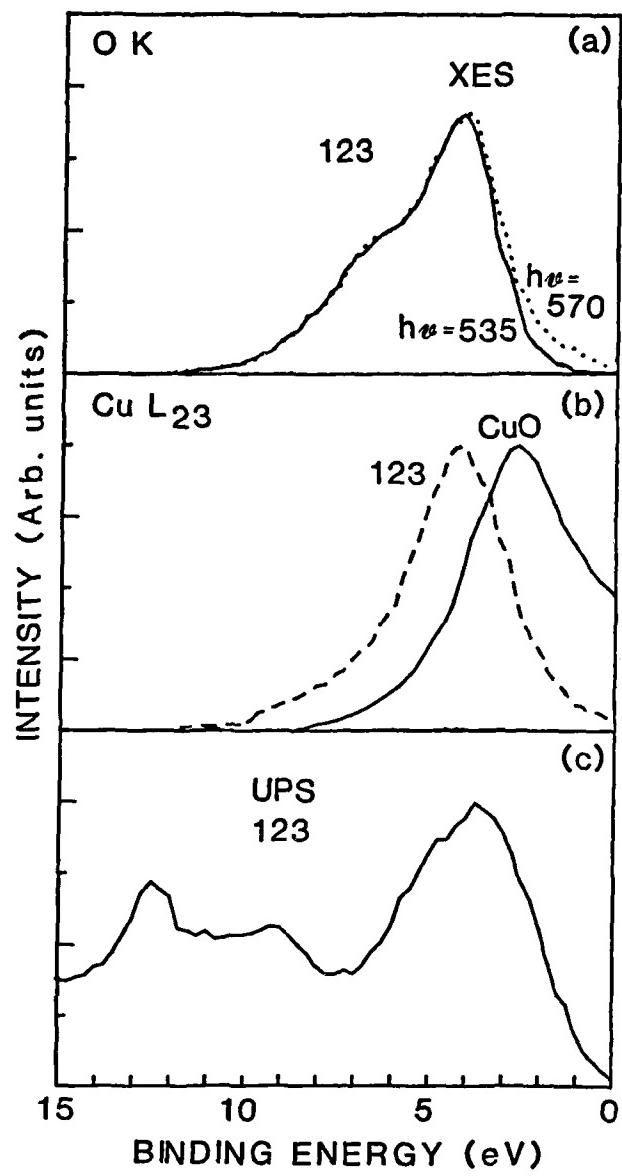
<sup>c</sup>The Calc. E and Exp. E columns indicate the results for 123.

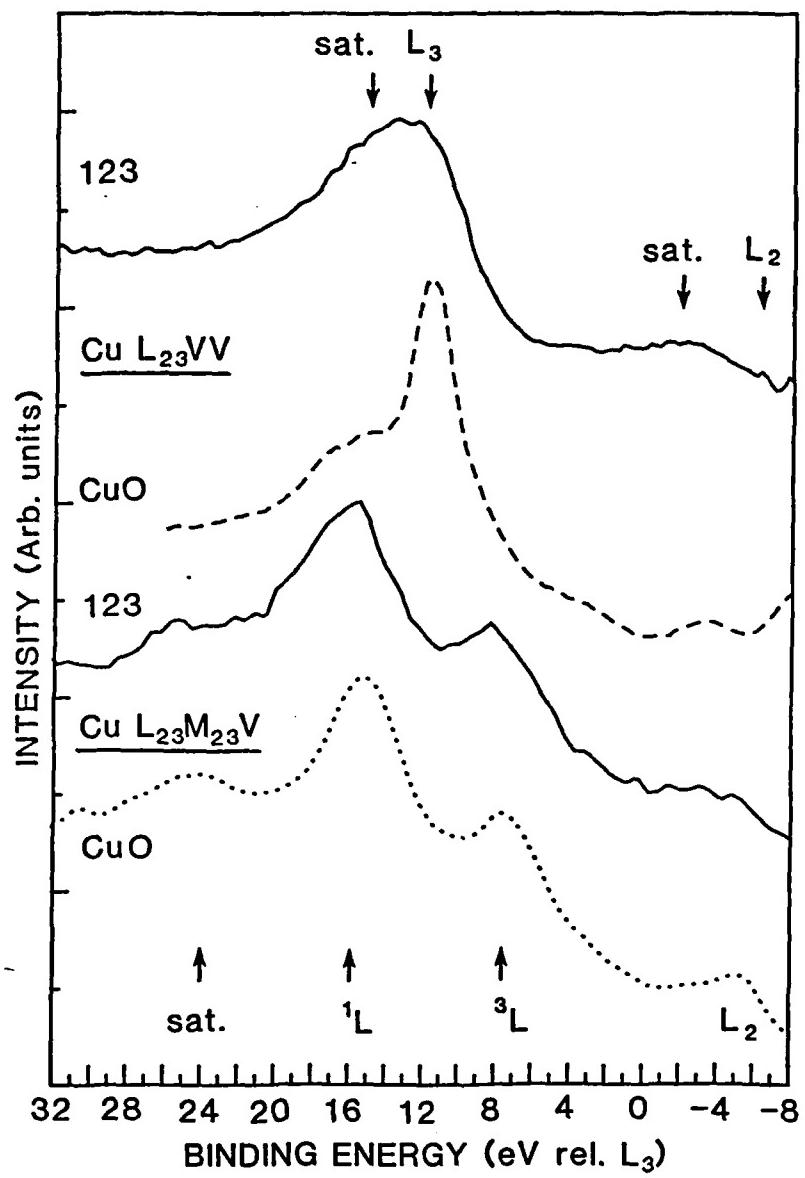
<sup>d</sup>The calculated E is defined relative to the ground v<sup>f</sup> (d) state energy =  $c_d - \alpha$ . The v<sup>f</sup>(d) energy defines the Fermi level relative to the vacuum level at zero.

<sup>e</sup>The dominant character switches as described in the text, and thus the sign in front of  $\delta_2$  is the opposite for CuO.

**Figure Captions**

- Fig. 1a)** Comparison of O K XES data for 123 taken at the indicated photon excitation energies (from Ref. 2).
- 1b)** Comparison of Cu L<sub>2,3</sub> XES data for 123 (Ref. 2) and CuO (Ref. 20).
- 1c)** UPS data for 123 ( $\hbar\nu \approx 74$  eV from Ref. 10).
- Fig. 2)** Comparison of Auger data for the materials indicated. Cu L<sub>2,3</sub>VV data for CuO and 123 from refs. 25 and 1. Cu L<sub>2,3</sub>M<sub>2,3</sub>V data for CuO from ref. 24 and for 123, this work. The L<sub>2,3</sub>VV data is on a 2-hole binding energy scale = E<sub>L3</sub>-E<sub>K</sub>, and the L<sub>2,3</sub>M<sub>2,3</sub>V on a 1-hole scale = E<sub>L3</sub>-E<sub>K</sub>-E<sub>M3</sub>, where E<sub>L3</sub> ≈ 933.4 and E<sub>M3</sub> ≈ 77.3 eV [9,11].





#### References

1. D.E. Ramaker et al., *Phys. Rev.* **36**, 5672 (1987).
2. K.L. Tsang et al., *Phys. Rev.* **B37**, 2293 (1988).
3. G.vanderLaan et al., *Phys. Rev.* **24**, 4369 (1981); J.C. Fuggle et al., *Phys. Rev.* **B37**, 1123 (1988).
4. J.E. Greidan et al., *Phys. Rev.* **B35**, 8770 (1987).
5. R.A. de Groot, H. Gutfreund, and M. Weger, *Sol. State Commun.* **63**, 451 (1987); W. Folkerts et al., *J. Phys. C: Solid State Phys.* **20**, 4135 (1987); A. Manthiram, X.X. Tang, and J.B. Goodenough, *Phys. Rev.* **B37**, 3734 (1988).
6. J.E. Hirsch et al., *Phys. Rev. Letters* **60**, 1168 (1988).
7. J. Redinger et al., *Phys. Lett.* **124**, 463 and 469 (1987).
8. D.E. Ramaker, N.H. Turner, and F.L. Hutson, submitted.
9. M.R. Thuler, R.L. Benbow, and Z. Hurvch, *Phys. Rev.* **B26**, 669 (1982).
10. N.G. Stoffel et al., *Phys. Rev.* **B37**, 7952 (1988); also preprint.
11. C. Benndorf et al., *J. Electron. Spectrosc. Related Phenom.* **19**, 77 (1980).
12. R. Kurtz et al., *Phys. Rev.* **B35**, 8818 (1987).
13. A. Rosencwaig and G.K. Wertheim, *J. Elect. Spectrosc. Related Phenom.* **1**, 493 (1972/73).
14. P. Steiner et al., *Z. Phys. B- Condensed Matter* **67**, 497 (1987).
15. D.E. Ramaker et al., *Phys. Rev.* **36**, 5672 (1987).
16. D.D. Sarma, *Phys. Rev.* **B37**, 7948 (1988).
17. S.L. Qiu et al., *Phys. Rev.* **B37**, 3747 (1988); W.K. Ford et al., *Phys. Rev.* **B37**, 7924 (1988); D.E. Ramaker, N.H. Turner, and F.L. Hutson, In Thin Film Processing and Characterization of High Temperature

Superconductors, J.M.Harper, J.H. Colton, and L.C. Feldman eds.,  
AVS Series No. 3 (American Institute Physics, New York, 1988), p  
284.

18. T. Takahashi et al., Phys. Rev. B37, 9788 (1988).
19. J.W. Gadzuk and M. Sunjic, Phys. Rev. B12, 524 (1975).
20. A.S. Koster, Mole. Phys. 26, 625 (1973).
21. D. van der Marel et al., Phys. Rev. B37, 5136 (1988).
22. Y. Chang et al., Phys. Rev. B (In press).
23. N. Nucker et al., Z. Phys. B: Cond. Matter 67, 9 (1987); Phys. Rev. 37, 5158 (1988).
24. L. Fiermans, R. Hoogewijs, and J. Vennik, Surf. Sci. 47, 1 (1975).
25. P.E. Larson, J. Electron Spectrosc. Related Phenom. 4, 213 (1974).

DL/1113/87/2

TECHNICAL REPORT DISTRIBUTION LIST, GEN

	No. Copies		No. Copies
Office of Naval Research Attn: Code 1113 800 N. Quincy Street Arlington, Virginia 22217-5000	2	Dr. David Young Code 334 NORDA NSTL, Mississippi 39529	1
Dr. Bernard Douda Naval Weapons Support Center Code 50C Crane, Indiana 47522-5050	1	Naval Weapons Center Attn: Dr. Ron Atkins Chemistry Division China Lake, California 93555	1
Naval Civil Engineering Laboratory Attn: Dr. R. W. Drisko, Code L52 Port Hueneme, California 93401	1	Scientific Advisor Commandant of the Marine Corps Code RD-1 Washington, D.C. 20380	1
Defense Technical Information Center Building 5, Cameron Station Alexandria, Virginia 22314	12 high quality	U.S. Army Research Office Attn: CRD-AA-IP P.O. Box 12211 Research Triangle Park, NC 27709	1
DTNSRDC Attn: Dr. H. Singerman Applied Chemistry Division Annapolis, Maryland 21401	1	Mr. John Boyle Materials Branch Naval Ship Engineering Center Philadelphia, Pennsylvania 19112	1
Dr. William Tolles Superintendent Chemistry Division, Code 6100 Naval Research Laboratory Washington, D.C. 20375-5000	1	Naval Ocean Systems Center Attn: Dr. S. Yamamoto Marine Sciences Division San Diego, California 91232	1

DL/1113/87/2

ABSTRACTS DISTRIBUTION LIST, 056/625/629

- Dr. J. E. Jensen  
Hughes Research Laboratory  
3011 Malibu Canyon Road -  
Malibu, California 90265
- Dr. J. H. Weaver  
Department of Chemical Engineering  
and Materials Science  
University of Minnesota  
Minneapolis, Minnesota 55455
- Dr. A. Reisman  
Microelectronics Center of North Carolina  
Research Triangle Park, North Carolina  
27709
- Dr. M. Grunze  
Laboratory for Surface Science  
and Technology  
University of Maine  
Orono, Maine 04469
- Dr. J. Butler  
Naval Research Laboratory  
Code 6115  
Washington D.C. 20375-5000
- Dr. L. Interante  
Chemistry Department  
Rensselaer Polytechnic Institute  
Troy, New York 12181
- Dr. Irvin Heard  
Chemistry and Physics Department  
Lincoln University  
Lincoln University, Pennsylvania 19352
- Dr. K. J. Klabunde  
Department of Chemistry  
Kansas State University  
Manhattan, Kansas 66506
- Dr. C. B. Harris  
Department of Chemistry  
University of California  
Berkeley, California 94720
- Dr. R. Bruce King  
Department of Chemistry  
University of Georgia  
Athens, Georgia 30602
- Dr. R. Reeves  
Chemistry Department  
Rensselaer Polytechnic Institute  
Troy, New York 12181
- Dr. Steven M. George  
Stanford University  
Department of Chemistry  
Stanford, CA 94305
- Dr. Mark Johnson  
Yale University  
Department of Chemistry  
New Haven, CT 06511-8118
- Dr. W. Knauer  
Hughes Research Laboratory  
3011 Malibu Canyon Road  
Malibu, California 90265
- Dr. Theodore E. Madey  
Surface Chemistry Section  
Department of Commerce  
National Bureau of Standards  
Washington, D.C. 20234
- Dr. J. E. Demuth  
IBM Corporation  
Thomas J. Watson Research Center  
P.O. Box 218  
Yorktown Heights, New York 10598
- Dr. M. G. Lagally  
Department of Metallurgical  
and Mining Engineering  
University of Wisconsin  
Madison, Wisconsin 53706
- Dr. R. P. Van Duyne  
Chemistry Department  
Northwestern University  
Evanston, Illinois 60637
- Dr. J. M. White  
Department of Chemistry  
University of Texas  
Austin, Texas 78712
- Dr. Richard J. Saykally  
Department of Chemistry  
University of California  
Berkeley, California 94720

DL/1113/87/2

ABSTRACTS DISTRIBUTION LIST, 056/625/629

Dr. F. Carter  
Code 6170  
Naval Research Laboratory  
Washington, D.C. 20375-5000

Dr. Richard Colton  
Code 6170  
Naval Research Laboratory  
Washington, D.C. 20375-5000

Dr. Dan Pierce  
National Bureau of Standards  
Optical Physics Division  
Washington, D.C. 20234

Dr. R. G. Wallis  
Department of Physics  
University of California  
Irvine, California 92664

Dr. D. Ramaker  
Chemistry Department  
George Washington University  
Washington, D.C. 20052

Dr. J. C. Hemminger  
Chemistry Department  
University of California  
Irvine, California 92717

Dr. T. F. George  
Chemistry Department  
University of Rochester  
Rochester, New York 14627

Dr. G. Rubloff  
IBM  
Thomas J. Watson Research Center  
P.O. Box 218  
Yorktown Heights, New York 10598

Dr. J. Baldeschwieler  
Department of Chemistry and  
Chemical Engineering  
California Institute of Technology  
Pasadena, California 91125

Dr. Galen D. Stucky  
Chemistry Department  
University of California  
Santa Barbara, CA 93106

Dr. A. Steckl  
Department of Electrical and  
Systems Engineering  
Rensselaer Polytechnic Institute  
Troy, New York 12181

Dr. John T. Yates  
Department of Chemistry  
University of Pittsburgh  
Pittsburgh, Pennsylvania 15260

Dr. R. Stanley Williams  
Department of Chemistry  
University of California  
Los Angeles, California 90024

Dr. R. P. Messmer  
Materials Characterization Lab.  
General Electric Company  
Schenectady, New York 22217

Dr. J. T. Keiser  
Department of Chemistry  
University of Richmond  
Richmond, Virginia 23173

Dr. R. W. Plummer  
Department of Physics  
University of Pennsylvania  
Philadelphia, Pennsylvania 19104

Dr. E. Yeager  
Department of Chemistry  
Case Western Reserve University  
Cleveland, Ohio 41106

Dr. N. Winograd  
Department of Chemistry  
Pennsylvania State University  
University Park, Pennsylvania 16802

Dr. Roald Hoffmann  
Department of Chemistry  
Cornell University  
Ithaca, New York 14853

Dr. Robert L. Whetten  
Department of Chemistry  
University of California  
Los Angeles, CA 90024

Dr. Daniel M. Neumark  
Department of Chemistry  
University of California  
Berkeley, CA 94720

Dr. G. H. Morrison  
Department of Chemistry  
Cornell University  
Ithaca, New York 14853

DL/1113/87/2

ABSTRACTS DISTRIBUTION LIST, 056/625/629

Dr. G. A. Somorjai  
Department of Chemistry  
University of California  
Berkeley, California 94720

Dr. J. Murday  
Naval Research Laboratory  
Code 6170  
Washington, D.C. 20375-5000

Dr. W. T. Peria  
Electrical Engineering Department  
University of Minnesota  
Minneapolis, Minnesota 55455

Dr. Keith H. Johnson  
Department of Metallurgy and  
Materials Science  
Massachusetts Institute of Technology  
Cambridge, Massachusetts 02139

Dr. S. Sibener  
Department of Chemistry  
James Franck Institute  
5640 Ellis Avenue  
Chicago, Illinois 60637

Dr. Arnold Green  
Quantum Surface Dynamics Branch  
Code 3817  
Naval Weapons Center  
China Lake, California 93555

Dr. A. Wold  
Department of Chemistry  
Brown University  
Providence, Rhode Island 02912

Dr. S. L. Bernasek  
Department of Chemistry  
Princeton University  
Princeton, New Jersey 08544

Dr. W. Kohn  
Department of Physics  
University of California, San Diego  
La Jolla, California 92037

Dr. Stephen D. Kevan  
Physics Department  
University Of Oregon  
Eugene, Oregon 97403

Dr. David M. Walba  
Department of Chemistry  
University of Colorado  
Boulder, CO 80309-0215

Dr. L. Kesmodel  
Department of Physics  
Indiana University  
Bloomington, Indiana 47403

Dr. K. C. Janda  
University of Pittsburgh  
Chemistry Building  
Pittsburgh, PA 15260

Dr. E. A. Irene  
Department of Chemistry  
University of North Carolina  
Chapel Hill, North Carolina 27514

Dr. Adam Heller  
Bell Laboratories  
Murray Hill, New Jersey 07974

Dr. Martin Fleischmann  
Department of Chemistry  
University of Southampton  
Southampton SO9 5NH  
UNITED KINGDOM

Dr. H. Tachikawa  
Chemistry Department  
Jackson State University  
Jackson, Mississippi 39217

Dr. John W. Wilkins  
Cornell University  
Laboratory of Atomic and  
Solid State Physics  
Ithaca, New York 14853

Dr. Ronald Lee  
R301  
Naval Surface Weapons Center  
White Oak  
Silver Spring, Maryland 20910

Dr. Robert Gomer  
Department of Chemistry  
James Franck Institute  
5640 Ellis Avenue  
Chicago, Illinois 60637

Dr. Horia Metiu  
Chemistry Department  
University of California  
Santa Barbara, California 93106

Dr. W. Goddard  
Department of Chemistry and Chemical  
Engineering  
California Institute of Technology  
Pasadena, California 91125